

Amendments to the Specification:

Please amend paragraphs [0008], [0024], [0025], and [0029] to read as follows:

[0008] An oil-based emulsion gravel-pack carrier fluid was also reported in SPE Paper 64978 that further discusses the possibility of adding a chelating agent solution to the internal phase that is for instance released by breaking the emulsion. This paper further depicts the effect of mixing energy on emulsion viscosity and the sand carrying properties of this oil-based emulsion fluid. However this paper is silent as to the choice of the emulsifier and the authors of the present invention have found that proper selection of said emulsifier was indeed critical for providing an effective oil-based emulsion gravel-pack carrier fluid. It is well known that commercial emulsifiers such as fatty acid esters are often blends of the mono, di, tri etc. esters, and also may be prepared under uncontrolled or different reaction conditions with differing temperatures, reactants, catalysts, etc. Thus, there is a need for a way to predict or determine whether or not a particular emulsifier will be effective in an oil-based emulsion gravel-pack carrier fluid.

[0024] Figure 1 shows two types of responses between about 22 and about 28 minutes. In both cases, a series of peaks are observed between about 22 and about 26 minutes which are depicted to correspond to the dialkylester, the trialkylester and tetraalkylester which are also referred to in the art as the dimer and trimer and tetramer. The last peak to be observed in that period of time is a little before 28 minutes and is depicted to be the low molecular weight or the monomer (monoester) peak. Note that depending on the chromatography equipment, and on the nature and weight of the polyester polymers, the peaks may appear slightly earlier or later, in

other words, all peaks are shifted towards earlier or later time, but the monomer (monoester) peak remains always present. With the first surfactant (open circles), no other peak can be detected before the peak 1 between about 26 and 28 minutes. In the second case on the other hand (full squares), a shoulder peak is clearly identified (indeed, the shoulder peak 3 is greater than the monomer peak 2).

[0025] Figure 2 shows a detail of the graph near the monomer peak and just before, in other words during the period of time just before the peak attributed to the sorbitan monomer (ie monooleate). Some of the tested surfactants clearly exhibit a shoulder peak. This is in particular the case with #1 commercialized as ~~sodium~~ sorbitan monooleate. Surfactants #2 and #3 also exhibit a shoulder peak, even if smaller than #1. The last tested surfactants, #4, #5 and #6 exhibit virtually no response during that period of time. To be noted that #6 is commercialized as ~~sodium~~ sorbitan trioleate.

[0029] In the last series of experiments, the stability of the emulsion after a medium (1 day) time and long time (9 days) was tested. The emulsions were made with different variant of commercial sorbitan oleate esters having different peaks. Figure 5 shows the percentage of volume broke out in 1 and 9 days at 150°F. Again, the tests show that the higher the shoulder peak, the higher the stability. Indeed, the suspension based on "pure" monooleate sorbitan (the chromatogram shows the presence of other fatty acid ester species) may be considered too stable and causing formation damage. Therefore, the suspension having a peak between 4 and 2 may be more suitable.